

Article

Tellurium Minerals: Structural and Chemical Diversity and Complexity

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Abstract: The chemical diversity and complexity of tellurium minerals were analyzed using the concept of mineral systems and Shannon informational entropy. The study employed data for 176 Te mineral species known today. Tellurium minerals belong to six mineral systems in the range of one-to-six species-defining elements. For 176 tellurium minerals, only 36 chemical elements act as essential species-defining constituents. The numbers of minerals of main elements are calculated as follows (the number of mineral species is given in parentheses): O (89), H (48), Cu (48), Pb (43), Bi (31), S (29), Ag (20), Fe (20), Pd (16), Cl (13), and Zn (11). In accordance with their chemistry, all Te minerals are classified into five types of mineral systems: tellurium, oxides, tellurides and intermetalides, tellurites, and tellurates. A statistical analysis showed positive relationships between the chemical, structural, and crystallochemical complexities and the number of essential species-defining elements in a mineral. A positive statistically significant relationship between chemical and structural complexities was established. It is shown that oxygen-free and oxygen-bearing Te minerals differ sharply from each other in terms of chemical and structural complexity, with the first group of minerals being simpler than the second group. The oxygen-free Te minerals (tellurium, tellurides, and intermetallides) are formed under reducing conditions with the participation of hydrothermal solutions. The most structurally complex oxygen-bearing Te minerals originate either from chemical weathering and the oxidation of ore deposits or from volcanic exhalations (Nabokoite).

Keywords: tellurium minerals; structural complexity; chemical complexity; crystallochemical complexity; mineral systems; Shannon informational entropy

1. Introduction

Mineral ecology is a branch of mineralogy that investigates the factors that affect the distribution of minerals and the evolution of their complexity and diversity in space and time. The concept of mineral ecology appeared recently [1], although some of the ideas underlying the basic principles of this approach were formulated and discussed by Russian mineralogists [2–6]. In this paradigm, studies on the mineral evolution of individual chemical elements are of particular interest [7–16]. Among the tasks of mineral ecology is to study the distribution patterns of minerals in the Earth's crust, their diversity and complexity, and their behavior in the human environment.

The emergence of new digital technologies for big-data analysis has revolutionized this field, leading to many important discoveries in the field of structural, chemical, and genetic relationships between various minerals, and it has formulated a new research direction based on these discoveries in mineralogy [17].

Yushkin [5] justified the expediency of formulating quantitative criteria for describing the state of the mineralogical kingdom at certain stages of its development over time, and Petrov [18],

Bulkin, [19] and Yushkin [20] pointed out that the informational entropy approach can be used to measure the chemical complexity and diversity of mineralogical and geochemical systems. S. Krivovichev [21–25] proposed to use Shannon information entropy for the quantitative analysis of the structural complexities of minerals and outlined the main applications of this approach to understanding the structural evolution of minerals. It was also shown that the parameters of structural complexity are directly related to the configurational entropy of crystalline solids [26]. V. Krivovichev and Charykova [27,28] and Grew et al. [12] proposed an alternative approach based on the concept of mineral systems, which provides a useful tool for the systematization of mineral species (e.g., [27]) and allows one to order existing data into coherent structures that emphasize changes in the diversity of minerals and their composition over time. In addition, it was shown that using the concept of mineral systems, it is possible to compare different geological objects with each other in terms of their mineral diversity [29–32]. Shannon information entropy (chemical and structural complexities) and the concepts of mineral systems have also been applied to mineral evolution as a whole [33–35] and to the mineral ecology of selenium minerals, including their diversity and complexities [36–38].

The main goal of this work was to study the mineral ecology of Te minerals. The tellurium chemistry resembles that of sulfur and selenium due to the proximity of these elements in the periodic table. Similarly to S and Se, tellurium adopts four oxidation states: -2 (tellurides), 0 (native tellurium), $+4$ (tellurites), and $+6$ (tellurates). It is generally assumed that the crystal structures and reactivity of tellurium compounds can be extrapolated from S and Se. In reality, recent discoveries and well-founded observations have shown that this is not entirely true. This was shown in a special review of the fundamental concepts that are necessary to understand the unique features of Te chemistry, with an emphasis on the difference between sulfur and selenium [39]. Tellurium minerals are relatively rare, and, currently, only 176 mineral species are known [40] (Supplementary Table S1). In minerals, Te participates in various forms of intermetallic bonding in tellurides [41–43], e.g., in hessite and calaverite [44]. In O-bearing species, tellurium occurs as Te^{4+} and/or Te^{6+} cations. The geometric configuration of Te–O bonds around Te^{m+} centres form $[\text{Te}^{m+}\text{O}_n]^{m-2n}$ oxyanions [44,45]. Tetravalent Te has a stereochemically active lone pair of electrons and usually possesses an asymmetrically distorted coordination polyhedra, e.g., $[\text{Te}^{4+}\text{O}_3]^{2-}$, $[\text{Te}^{4+}\text{O}_4]^{4-}$, and $[\text{Te}^{4+}\text{O}_5]^{6-}$, sometimes with more than one coordination in the same structure [44,46]. On the contrary, Te^{6+} lacks a lone electron pair and occurs in minerals exclusively as $[\text{Te}^{6+}\text{O}_6]^{6-}$ octahedra. However, natural tellurates are stable under oxidizing conditions and, unlike selenates, do not transform into tellurites, which explains the relatively large number of tellurates in nature. The structural diversity of O-bearing Te minerals is further discussed below.

It is interesting to compare the mineral and atomic abundances in the Earth's crust for tellurium and selenium. Based on the data listed in [47], the crustal abundance by atoms for Te is less than that for Se by approximately 80 times (0.00016 and 0.013 ppm, respectively). However, the number of Te minerals exceeds that of selenium by 1.4 times (176 and 123, respectively). Thus, with a general positive trend of an increasing number of mineral species of the same element with an increase in its crustal abundance by atoms [45], tellurium differs from selenium by an anomalous mineral diversity. This difference between the atomic and mineral abundances of Te and Se is due to several reasons. Under endogenic conditions, tellurium, in contrast to selenium, is not disseminated in the sulfides and selenides and tends to form its own mineral species [41]. In addition, a certain contribution to the increase in tellurium minerals is also made by its ability to form intermetallic compounds under reducing conditions (e.g., bilibinskite, bezsmertnovite, and bogdanovite) [41].

The different behaviors of tellurium and selenium are manifested to an even greater extent by the comparison of the number of oxygen-bearing minerals (Se, 34; Te, 86) that are stable under oxidizing conditions. The latter is due to the uncommon diversity of the crystal structures of oxygen-bearing tellurium compounds [48–50], namely: the ability of Te O_n polyhedra to polymerize, forming anionic groups in the form of oligomers, chains, layers, and complex three-dimensional

frameworks. Te^{6+} forms $[\text{Te}^{6+}\text{O}_6]^{6-}$ octahedra only, whereas Te^{4+} has a lone electron pair and can form several types of asymmetrical $[\text{Te}^{4+}\text{O}_n]$ polyhedra [51].

Earlier, we noted [52] a substantial difference of sulfur and selenium behavior in the near-surface oxidizing settings: Sulfur here is encountered either in a native state or in the sulfate form (oxidation state +6), very rarely in sulfite form (oxidation state +4). Conversely, selenium is encountered more often as selenites (oxidation state +4), while selenates (oxidation state +6) are quite rare with only three known mineral species. Tellurium minerals, as mentioned above, exist under oxidizing conditions as both tellurites and tellurates. Thus, selenium differs from sulfur and tellurium, which are in the VI group of the periodic table, by a diminished stability of natural compounds in the highest oxidation state. It is known from the chemical literature that this feature is generally characteristic for the chemical elements directly following elements in the periodic table with filled sublevels of a certain type (*p*, *d*, and *f*). This category includes, in particular, *p* elements of period IV (As, Se, Br, etc.)—in their electron configurations, the inner $3d^{10}$ electron shell first appears, which is absent in the elements of the preceding periods [53,54].

Previously, we have characterized selenium minerals and described their chemical and structural complexities [36–38]. The aim of this paper was to characterize Te minerals and to determine relations between their chemical, structural, and crystallochemical complexities, as well as to apply the data obtained to the understanding of Te mineral complexity and diversity.

2. Materials and Methods

2.1. Mineral Systems

There are only 176 Te mineral species known to date [39, mindat.org]. It has been shown previously [27,28] that any mineral can be assigned to a mineral system according to species-defining elements in its chemical formula. These species-defining elements were chosen according to the IMA (International Mineralogical Association) rules for the definition of a new mineral species [55–60], taking into account the discussion of some problems in Russian literature [61].

The chemical formulae of Te minerals used to compose mineral systems are taken from constantly updated lists published by Marco Pasero [39] at the website of Commission on New Minerals, Nomenclature, and Classification IMA (CNMNC IMA). We identified the mineral system for each tellurium mineral of the set of species-defining elements; e.g., for the above-mentioned goldfieldite, this is the three-component system STeCu. When attributing a mineral to some system, we only took the species-defining elements into account, ignoring all other isomorphous components, even geochemically important ones [27,28]. A chemical system is identified in accordance with the thermochemical sequence of chemical elements (Figure 1) [27]. For example, yecoraite, $\text{Fe}^{3+}_3\text{Bi}_5\text{O}_9(\text{Te}^{4+}\text{O}_3)(\text{Te}^{6+}\text{O}_4)_2 \cdot 9\text{H}_2\text{O}$, belongs to the system OHTeBiFe, while tsnigriite, $\text{Ag}_9\text{SbTe}_3\text{S}_3$, belongs to the system STeSbAg.

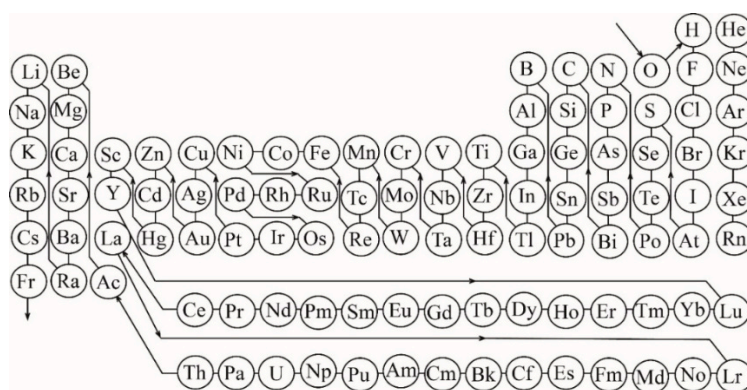


Figure 1. Thermochemical sequence of chemical elements. Arrows indicate the order of chemical elements used to identify a mineral system.

2.2. Chemical, Structural and Crystallochemical Complexities

To quantify the chemical, structural and crystallochemical complexities of Te minerals, a total of pieces of 176 chemical composition data, and 117 data sets on crystal structure were reviewed. The level of structural complexity per atom ($^{str}I_G$) and per unit cell ($^{str}I_{G,total}$) [21–26] were calculated using the following equations [36]:

$$^{str}I_G = - \sum_{i=1}^k p_i \log_2 p_i \text{ (bits/atom)} \quad (1)$$

$$^{str}I_{G,total} = -v \sum_{i=1}^k p_i \log_2 p_i \text{ (bits/cell)} \quad (2)$$

where k is the number of different crystallographic orbits (Wyckoff sites) in the structure and p_i is the random choice probability for an atom from the i th crystallographic orbit, that is:

$$p_i = m_i/v \quad (3)$$

where m_i is a multiplicity of a crystallographic orbit (i.e., the number of atoms of a specific Wyckoff site in the reduced unit cell) and v is the total number of atoms in the reduced unit cell.

For several hydrated crystal structures of Te minerals, the proton positions have not been determined. In these cases, the H-correction procedure has been applied [62]. All structure complexity calculations have been performed by means of the TOPOS program package [63].

The chemical complexities of tellurium minerals were evaluated by the amount of chemical information per atom ($^{chem}I_G$) and per formula unit, f.u. ($^{chem}I_{G,total}$) [35,36]. The chemical information was calculated as follows [36]:

$$^{chem}I_G = - \sum_{i=1}^k p_i \log_2 p_i \text{ (bits/atom)} \quad (4)$$

$$^{chem}I_{G,total} = -e \sum_{i=1}^k p_i \log_2 p_i \text{ (bits/cell)} \quad (5)$$

where k is the number of species-defining chemical elements and p_i is the random choice probability for an atom of the i th element, that is:

$$p_i = c_i/e \quad (6)$$

where e is the total number of species-defining elements in the chemical formula:

$$e = \sum_{i=1}^k c_i \quad (7)$$

For the calculations of chemical complexities, the ideal chemical formulas of minerals were used (see also [27,28]).

Additionally, we calculated the crystal chemical complexities ($^{chem+str}I_G$ and $^{chem+str}I_{G,total}$) as the sum of chemical and structural complexities:

$$^{chem+str}I_G = ^{chem}I_G + ^{str}I_G \text{ (bits/atom)} \quad (8)$$

$$^{chem+str}I_{G,total} = ^{chem}I_{G,total} + ^{str}I_{G,total} \text{ (bits/f.u., cell)} \quad (9)$$

3. Results

3.1. Classification of Tellurium Minerals

Our approach allowed us to systematize mineral species by their chemical composition and to organize the available data in such a way that they complemented traditional classification schemes [61,64,65]. A certain advantage of our approach to mineral classification is its formal and unambiguous nature, although we are aware that some significant information was lost here.

As in selenium minerals classification [36], tellurium minerals are divided into five groups: native tellurium, oxides, tellurides and intermetallides, tellurites, and tellurates (anhydrous, hydrous, and with additional anions). In each of these groups, minerals can be classified by mineral systems (or the minimum number of species-forming elements). In accordance with this value, six Te mineral systems can be distinguished, with them containing from one to six species-defining elements. Native tellurium belongs to the one-component mineral system, while minerals consisting of two elements (altaite, kotulskite, hessite, etc.) belong to two-component systems, minerals containing three elements (petzite, monchetundraite, sopcheite, etc.) belong to three-component systems, etc. Our approach allowed us to distinguish the formulae of “chemically pure” minerals that contain only species-defining elements [27,28,33,34]. In this system, each mineral occupies a unique position determined by the number of species-defining elements. This also facilitates the use of digital technologies for organizing, storing, and retrieving thermodynamic data for a particular mineral [36–38]. The classification of the Te mineral systems is given in Supplementary Table S2.

3.2. Distribution of Te Minerals in Accordance with the Number of Species-Defining Elements

Tellurium minerals formed mainly in hydrothermal deposits (e.g., Bonanza Mining District, Colorado, USA [66]; Moctezuma mines, Mexico [67,68]; Alekseevskoye Mine, Sakha Republic, Russia [69]; Aginskoe deposit and Kamchatka Oblast', Russia [70,71]), the oxidation zones of ore mineral deposits (e.g., Otto Mountain, USA [49,50]; and Tombstone, Arizona, USA [72]) and very rare as products of volcanic fumaroles [73]. Tellurides are also found in magmatic Cu–Ni–PGM (platinum-group-metal) sulfide deposits (e.g., Noril'sk [74] and Kola Peninsula [75] in Russia; Bushveld in South Africa [76]; and Sudbury in Canada [77]), where they have a hydrothermal origin.

Similarly to the selenium minerals [36–38], we divided tellurium minerals into two groups based on their mineral composition: oxygen-free (native tellurium, tellurium intermetallides, and tellurides; 86 minerals in total) and oxygen-bearing (tellurium oxides, tellurites, and tellurates; 89 minerals, in total). Oxygen-free tellurium minerals are quite well studied and widespread in endogenic deposits of various genetic types. In rare sulfosalts, tellurium is a crystal chemical analog of As (e.g., in goldfieldite $\text{Cu}_{10}(\text{TeS}_3)_4\text{S}$, a mineral of the tetrahedrite group; see [78,79]). In near-surface settings, tellurium (similarly to chalcophile elements) is easily oxidized, forming oxides and oxysalts when it is in an oxidation state of +4 and/or +6.

The data are given in Table 1 and Figure 2. The arithmetic mean of species-defining elements in O-free Te minerals was significantly lower ($\bar{X} = 2.8$; $\sigma_{\bar{X}} = 0.09$) than in O-bearing minerals ($\bar{X} = 4.4$; $\sigma_{\bar{X}} = 0.11$). According to Student's test, the differences in the number of species-forming elements for both groups of Te minerals were statistically significant (the confidence level is more than 99.99%)

Table 1. The mineral systems (N) of Te minerals.

N	All Minerals		O-Free		O-Bearing	
	m_i	p_i	m_i	p_i	m_i	p_i
1	1	0.6	1	1.2	-	-
2	33	18.7	31	35.6	2	2.3
3	61	34.7	42	48.3	19	21.3
4	35	29.9	8	9.2	27	30.3
5	30	17.0	5	5.7	25	28.1
6	16	9.1	-	-	16	18.0
Total	176	100.0	87	100.0	89	100.0

Note: N —the mineral systems (the number of the species-defining elements); m —number of minerals; $p_i = (m_i / \sum_{i=1}^6 m_i) \cdot 100$ probability, %.

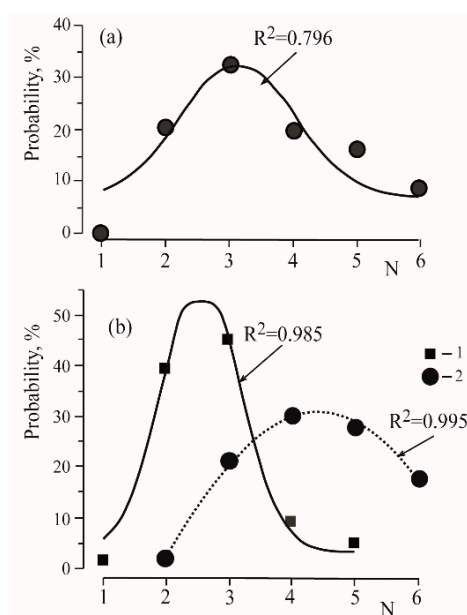


Figure 2. Distribution of Te minerals among mineral systems based on the number of the species-defining elements: (a) All minerals, (b) O-free (1, solid line), and O-bearing (2, dashed) minerals.

The distribution of all tellurium minerals versus the mineral systems (number of species-defining elements), N , was found to be close to normal (Figure 2a). The largest number of minerals was found to consist of two and three species-defining elements (36 and 58 minerals, respectively). It is worthy to note that all minerals containing six species-defining elements were found to be O-bearing. The distributions of Te minerals between O-free and O-bearing groups depending on the number of species-defining elements, N , were also normal (Table 1 and Figure 2b). The number of species-forming elements in minerals ranged from one to six, and the maximum number of minerals was found to be formed by two and three species-forming elements.

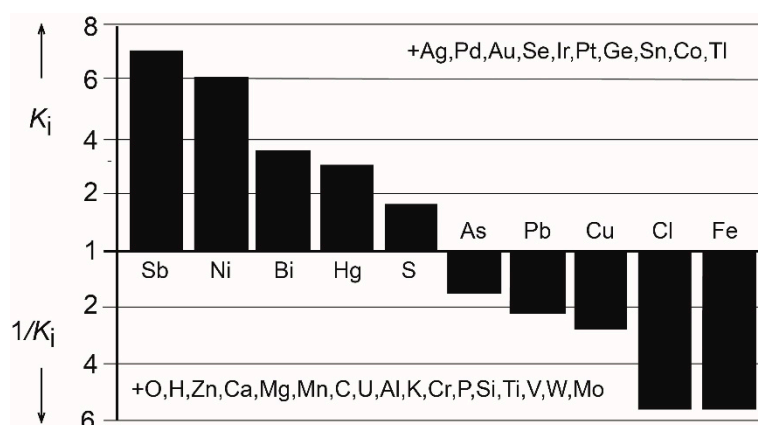
For tellurium minerals, 36 chemical elements were found to be species-defining elements. These elements were divided into three groups: (1) Elements that occur only in O-free minerals, (2) elements that occur in O-free and O-bearing minerals, and (3) elements that occur only in O-bearing tellurium minerals (Table 2).

Table 2. The species-defining elements of Te minerals separated into three groups: (1) All minerals, (2) O-free, and (3) O-bearing minerals.

Groups	Elements	All	O-Free	O-Bearing	* K_i
1	Ag	20	20	-	-
	Pd	16	16	-	-
	Au	13	13	-	-
	Se	8	8	-	-
	Ir, Pt	3	3, each	-	-
	Sn	2	2	-	-
	Co, Ge, Tl	1	1	-	-
2	Sb	8	7	1	7.00
	Ni	7	6	1	6.00
	Hg	5	4	1	4.00
	Bi	31	24	7	3.43
	S	29	18	11	1.64
	As	6	3	3	1.00
	Pb	43	14	29	0.48
	Cu	48	13	33	0.39
	Cl	13	2	11	0.18
	Fe	20	2	18	0.11
3	O	89	-	89	-
	H	48	-	48	-
	Zn	11	-	11	-
	Ca	8	-	8	-
	Mg	7	-	7	-
	Mn	6	-	6	-
	U	4	-	4	-
	C	4	-	4	-
	Al	2	-	2	-
	K	2	-	2	-
	P, Cr, Mo, Si, Ti, V, W	1	-	1, each	-

* $K_i = \text{O-free/O-bearing}$.

The difference between chemical compositions of O-free and O-bearing tellurium minerals is more clearly illustrated by the diagram shown in Figure 3.

**Figure 3.** The species-defining elements (N) of three groups of Te minerals and the concentration coefficient $K_i = \text{O-free/O-bearing}$ (see Table 2).

3.3. Chemical, Structural and Crystallochemical Complexities of Tellurium Minerals

According to Equations (1)–(9), the complexity indicators of tellurium minerals belonging to different mineral systems (or the number (N) of different species-defining chemical elements) were calculated. The average chemical, structural, and crystal–chemical complexities of minerals and associated statistical parameters are shown in Table 3 and Figure 4. The dependencies of different complexities from N were approximated by the following functions (the corresponding curves are plotted in Figure 3):

$$I_G^{\text{chem}} = 1.900 - 3.743 \times \exp(-N/1.472) \quad (R^2 = 0.997) \quad (10)$$

$$I_G^{\text{str}} = 3.580 - 7.239 \times \exp(-N/1.397) \quad (R^2 = 0.979) \quad (11)$$

$$I_{G,\text{total}}^{\text{chem}} = 1.622 \times \exp(N/1.459) \quad (R^2 = 0.999) \quad (12)$$

$$I_{G,\text{total}}^{\text{str}} = -54.512 + 53.873 \times N \quad (R^2 = 0.985) \quad (13)$$

$$I_G^{\text{chem+str}} = 5.520 - 11.089 \times \exp(-N/1.396) \quad (R^2 = 0.979) \quad (14)$$

$$I_{G,\text{total}}^{\text{chem+str}} = -80.254 + 69.784 \times N \quad (R^2 = 0.980) \quad (15)$$

Table 3. Chemical, structural, and crystallochemical (chem + str) complexities of tellurium minerals according to the number of species-defining chemical elements (N) *.

N	m	chem I_G				m_i	str I_G				chem+str I_G			
		[bits/atom]		[bits/f.u.]			[bits/atom]		[bits/cell]		[bits/atom]		[bits/f.u.cell]	
		\bar{X}	σ_X^-	\bar{X}	σ_X^-		\bar{X}	σ_X^-	\bar{X}	σ_X^-	\bar{X}	σ_X^-	\bar{X}	σ_X^-
1	1	0			0	29	1	0			0			
2	36	0.96	0.02	7.0	1.67	29	2.07	0.24	72.1	25.9	3.3	0.3	82.0	27.9
3	58	1.38	0.02	12.8	1.17	37	2.44	0.16	81.2	13.3	3.8	0.1	94.9	13.9
4	35	1.63	0.03	26.2	2.68	18	3.10	0.21	168.3	34.1	4.7	0.2	190.4	35.4
5	29	1.79	0.03	50.1	6.62	21	3.57	0.18	199.4	36.4	5.3	0.2	245.2	41.6
6	16	1.90	0.04	101.8	18.77	11	3.58	0.22	283.3	41.2	5.5	0.2	371.5	39.4

* m_i = number of minerals taken into account; \bar{X} = arithmetic mean; σ_X^- = standard error of arithmetic mean.

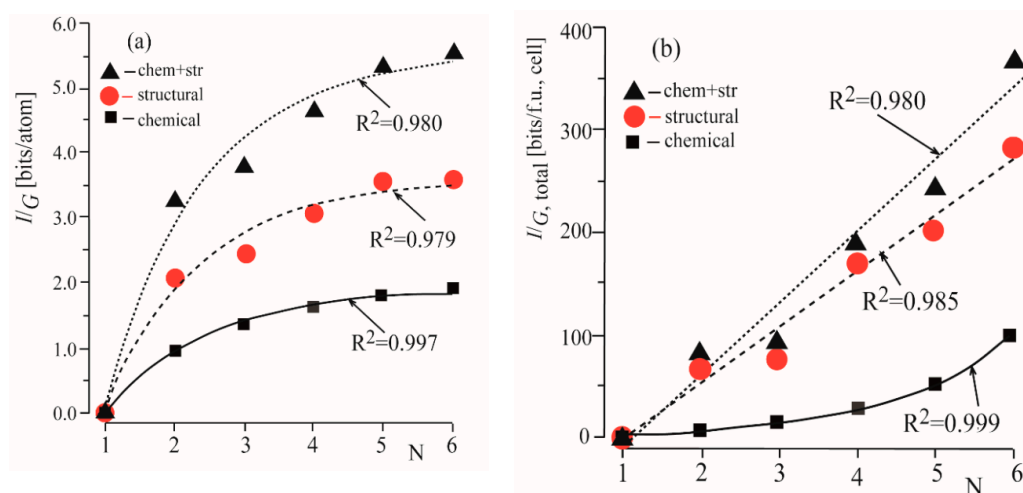


Figure 4. Dependence of chemical, structural, and crystal–chemical complexities of tellurium minerals on the number of species-defining chemical elements (N). The complexity per atom (a) and per unit cell or per formula unit (b).

The obtained data indicated the presence of a statistically significant (a confidence level of more than 99%) and positive relationship between the chemical, structural, and crystallochemical (chemical + structural) complexities of tellurium minerals and the number of species-defining chemical elements in their mineral formula.

To compare the complexity of tellurium minerals formed in various geochemical environments, they were divided into two groups: (1) oxygen-free minerals (tellurides and native tellurium) formed in endogenous conditions and (2) oxygen-bearing (oxides and oxysalts) formed in exogenous conditions as a result of chemical weathering and fumarolic activity. A statistical analysis showed (Table 4) that the average arithmetic values of structural, chemical, and crystallochemical complexities for O-free tellurium minerals are significantly lower than for O-bearing species. The Student's test [80] showed that the differences are statistically significant (see Table 4).

Table 4. Chemical, structural, and crystallochemical complexities of two groups (O-free and O-bearing) tellurium minerals *.

Complexities	O-free			O-bearing			Student's <i>t</i> -test	
	m_i	\bar{X}	$\sigma_{\bar{X}}$	m_i	\bar{X}	$\sigma_{\bar{X}}$	<i>t</i>	<i>p</i>
chem I_G [bits/atom]	87	1.28	0.05	89	1.62	0.03	−6.92	<0.0001
str I_G [bits/atom]	56	2.19	0.16	63	3.20	0.12	−5.10	<0.0001
chem $I_{G,\text{total}}$ [bits/f.u.]	87	11.4	1.1	89	45.9	5.2	−6.38	<0.0001
str $I_{G,\text{total}}$ [bits/cell]	56	69.3	15.6	63	183.9	18.5	−4.65	<0.0001
chem+str I_G [bits/atom]	56	3.69	0.20	63	4.82	0.14	−4.71	<0.0001
chem+str $I_{G,\text{total}}$ [bits/f.u.,cell]	56	81.2	16.6	63	223.8	21.1	−5.19	<0.0001

* m_i = number of minerals taken into account; \bar{X} = arithmetic mean; $\sigma_{\bar{X}}$ = standard error of arithmetic mean; *p*—probability value or *p*-value.

4. Conclusions

The obtained statistical correlations between the complexities (chem, str, and str + chem) of tellurium minerals showed that the minerals become more complex during the chemical weathering of primary ores. These indicators regularly and statistically significantly increase from the endogenic to exogenic associations, and they reflect the character and direction of mineral matter's variations in the evolution of mineral systems and can be used when analyzing their evolutionary patterns. This is in good agreement with the general trend of chemical differentiation of matter with the evolution of the Earth, which was pointed out in [1–5], where general patterns in the evolution of mineral species on Earth were established, including the complexity of the composition, structure, and diversity of minerals over geological time. In our case, this trend manifested itself as a complexity of later mineral assemblages compared to the primary ones. In addition, the average number of chemical elements in a mineral increases from primary to secondary minerals, which is consistent with the results of the information considerations presented here. On the example of selenium minerals [36], it was shown that the change in the complexity of minerals follows the same trend: more complex minerals occur over time, and simple ones are replaced. This trend of O-bearing minerals being more complex now seems an empirical rule that applies for many elements. That was found also for cobalt minerals [13].

Current work has shown that this conclusion is also true for the minerals of other elements, since the increase in their complexity over time reflects the general trend in the evolution of natural matter. Chemical differentiation leads to an increase in the complexity and diversity of tellurium minerals, which favor the formation of local concentrations of some rare elements and the creation of new geochemical environments.

In conclusion, let us note that our approach makes it possible to move from comparing the tellurium minerals of primary and secondary genesis from a qualitative to a quantitative basis. This gives a more vivid representation of the studied objects and can also reveal the ability of individual chemical elements to form their own mineral species.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2075-163X/10/7/623/s1>, Table S1: Tellurium minerals: Number of species-defining's elements (N) chemical formula, and number of localities (NL); Table S2: Classification of mineral systems of tellurium minerals according to number of species-defining elements (N) an chemical composition.

Author Contributions: Conceptualization, V.G.K. and S.V.K.; methodology, V.G.K. and S.V.K.; investigation, V.G.K., S.V.K. and M.V.C.; writing—original draft preparation, V.G.K.; writing—review and editing, V.G.K., S.V.K. and M.V.C. All authors have read and agreed to the published version of the manuscript.

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Table S1. Tellurium minerals: Number of species-defining's elements (*N*) chemical formula, and number of localities (*NL*).

<i>N</i>	Mineral	Chemical formula	<i>NL</i>	References
4	Adanite	Pb ₂ (TeO ₃)(SO ₄)	1	1
6	Agaitite	Pb ₃ [Cu(TeO ₅ OH)](CO ₃)(OH)	2	2
4	Alburnite	Ag ₈ GeTe ₂ S ₄	1	3
4	Aleksite	PbBi ₂ Te ₂ S ₂	28	4
2	Altaite	PbTe	489	5
5	Andychristyite	PbCu(TeO ₅)·H ₂ O	1	6
3	Andymcdonaldite	Fe ₂ TeO ₆	1	7
5	Backite	Pb ₂ AlTeO ₆ Cl	1	8
6	Bairdite	Pb ₂ Cu ₄ Te ₂ O ₁₀ (OH) ₂ (SO ₄)·H ₂ O	4	9
3	Baksanite	Bi ₆ Te ₂ S ₃	4	10
3	Balyakinite	Cu(TeO ₃)	4	11
5	Benleonardite	Ag ₁₅ CuSb ₂ S ₇ Te ₄	20	12
3	Bezsmertnovite	Au ₄ CuTe	6	13
4	Bilibinskite	PbAu ₃ Cu ₂ Te ₂	12	14
4	Bodieite	Bi ₂ (TeO ₃) ₂ (SO ₄)	4	15
3	Borovskite	Pd ₃ SbTe ₄	11	16
4	Brumadoite	Cu ₃ (TeO ₄)(OH) ₄ ·5H ₂ O	1	17
5	Buckhornite	Pb ₂ AuBiTe ₂ S ₃	13	18
6	Burckhardtite	Pb ₂ (Fe ³⁺ Te ⁶⁺) _{Σ2} (AlSi ₃ O ₈)O ₆	4	19
2	Calaverite	AuTe ₂	361	20
2	Cameronite	Cu ₅ Cu ₃ Te ₁₀	1	21
3	Carlfriesite	CaTe ⁶⁺ (Te ⁴⁺) ₂ O ₈	4	22
3	Cervelleite	Ag ₄ TeS	42	23
4	Cesbronite	Cu ₃ Te ⁶⁺ O ₄ (OH) ₄	3	24
3	Chekhovichite	Bi ₂ Te ₄ O ₁₁	5	25
4	Chenguodaite	Ag ₉ FeTe ₂ S ₄	2	26
5	Cheremnykhite	Pb ₃ Zn ₃ (TeO ₆)(VO ₄) ₂	1	27
4	Chiluite	Bi ₆ Te ₂ Mo ⁶⁺ ₂ O ₂₁	1	28
7	Choloalite	Pb ₃ (Cu ₂ Sb) _{Σ3} Te ₆ O ₁₈ Cl	7	29
5	Chromschieffelinite	Pb ₁₀ [Te ₂ O ₈ (OH) ₃] ₂ (TeO ₂ (OH) ₄) ₂ (CrO ₄)·5H ₂ O	1	30
3	Cliffordite	UTe ₃ O ₉	4	31
2	Coloradoite	HgTe	169	32
4	Cuzticite	Fe ₂ (TeO ₆)·3H ₂ O	2	33
3	Dagenaisite	Zn ₃ (TeO ₆)	1	34
5	Debattistiite	Ag ₁₈ HgAs ₁₂ S ₂₄ Te ₄	1	35
4	Denningite	CaMnTe ₄ O ₁₀	2	36
5	Dugganite	Pb ₃ Zn ₃ (TeO ₆)(AsO ₄) ₂	18	37
5	Eckhardtite	CaCu(TeO ₅) H ₂ O	3	38

4	Emmonsite	$\text{Fe}_2(\text{TeO}_3)_3 \cdot 2\text{H}_2\text{O}$	34	39
2	Empressite	AgTe	43	40
7	Eurekadumpite	$\text{Cu}_{16}(\text{TeO}_3)_2(\text{AsO}_4)_3\text{Cl}(\text{OH})_{18} \cdot 7\text{H}_2\text{O}$	5	41
5	Eztlite	$\text{Pb}_2\text{Fe}_3(\text{TeO}_3)_3(\text{SO}_4)\text{O}_2\text{Cl}$	2	33
3	Fairbankite	$\text{Pb}(\text{TeO}_3)$	1	42
4	Frankhawthorneite	$\text{Cu}_2(\text{TeO}_4)(\text{OH})_2$	2	43
2	Frohbergite	FeTe_2	40	44
6	Fuettererite	$\text{Pb}_3\text{Cu}_6(\text{TeO}_6)(\text{OH})_7\text{Cl}_5$	3	45
2	Gaotaiite	Ir_3Te_8	1	46
3	Goldfieldite	$\text{Cu}_{10}\text{Te}_4\text{S}_{13}$	72	47
4	Graemite	$\text{Cu}(\text{TeO}_3) \cdot \text{H}_2\text{O}$	6	48
6	Hagstromite	$\text{Pb}_8\text{Cu}(\text{TeO}_6)_2(\text{CO}_3)\text{Cl}_4$		49
2	Hedleyite	Bi_7Te_3	104	50
3	Henryite	$\text{Cu}_4\text{Ag}_3\text{Te}_4$	4	51
2	Hessite	Ag_2Te	804	52
4	Hitachiite	$\text{Pb}_5\text{Bi}_2\text{Te}_2\text{S}_6$	1	53
3	Honeaite	Au_3TlTe_2	3	54
5	Housleyite	$\text{Pb}_6\text{CuTe}_4\text{O}_{18}(\text{OH})_2$	5	55
6	Ilimneyite	$\text{MgZn}_2\text{Mn}_2(\text{TeO}_3)_6 \cdot 9\text{H}_2\text{O}$	1	56
3	Ingodite	Bi_2TeS	30	57
4	Jensenite	$\text{Cu}_3(\text{TeO}_6) \cdot 2\text{H}_2\text{O}$	3	58
6	Juabite	$\text{CaCu}_{10}(\text{TeO}_3)_4(\text{AsO}_4)_4(\text{OH})_2 \cdot 4\text{H}_2\text{O}$	6	59
2	Kalgoorlieite	As_2Te_3	2	60
3	Kawazulite	$\text{Bi}_2\text{Te}_2\text{Se}$	32	61
2	Keithconnite	$\text{Pd}_{20}\text{Te}_7$	40	62
7	Keystoneite	$\text{MgNi}_2\text{Fe}_2(\text{TeO}_3)_6 \cdot 9\text{H}_2\text{O}$	4	63
5	Khinite	$\text{Cu}_3\text{Pb}(\text{TeO}_6)(\text{OH})_2$	14	37
5	Kinichilite	$\text{MgMn}_2\text{Fe}_2(\text{TeO}_3)_6 \cdot 9\text{H}_2\text{O}$	2	64
3	Kitkaite	NiTeSe	1	65
3	Kochkarite	PbBi_4Te_7	5	66
3	Kojonenite	$\text{Pd}_{7-x}\text{SnTe}_2$	1	67
3	Kolarite	PbTeCl_2	3	68
3	Kostovite	AuCuTe_4	21	69
2	Kotulskite	PdTe	116	70
3	Krennerite	Au_3AgTe_8	111	71
5	Kuksite	$\text{Pb}_3\text{Zn}_3(\text{TeO}_6)(\text{PO}_4)_2$	6	27
4	Kuranakhite	$\text{PbMn}(\text{TeO}_6)$	7	72
3	Kurilite	$\text{Ag}_8\text{Te}_3\text{Se}$	2	73
5	Leisingite	$\text{CuMg}_2(\text{TeO}_6) \cdot 6\text{H}_2\text{O}$	2	74
2	Lingbaoite	AgTe_3	1	75
3	Lukkulaisvaaraite	$\text{Pd}_{14}\text{Ag}_2\text{Te}_9$	2	76

4	Mackayite	$\text{Fe}(\text{Te}_2\text{O}_5)(\text{OH})$	12	77
3	Magnolite	$\text{Hg}_2(\text{TeO}_3)$	3	78
3	Maletoyvayamite	$\text{Au}_3\text{Se}_4\text{Te}_6$	1	79
4	Markcooperite	$\text{Pb}_2(\text{UO}_2)(\text{TeO}_6)$	2	80
3	Maslovite	PtBiTe	23	81
2	Mattagamite	CoTe_2	7	82
3	Mayingite	IrBiTe	3	83
5	Mazzettiite	$\text{Ag}_3\text{HgPbSbTe}_5$	1	84
3	Mcalpineite	$\text{Cu}_3(\text{TeO}_6)$	12	85
2	Melonite	NiTe_2	207	86
2	Merenskyite	PdTe_2	155	87
5	Metatamboite	$\text{Fe}_3(\text{OH})(\text{SO}_4)(\text{TeO}_3)_3[\text{TeO}(\text{OH})_2] \cdot 3\text{H}_2\text{O}$	1	88
3	Michenerite	PdBiTe	114	89
3	Miessiite	$\text{Pd}_{11}\text{Te}_2\text{Se}_2$	1	90
4	Millsite	$\text{Cu}(\text{TeO}_3) \cdot 2\text{H}_2\text{O}$	1	91
2	Mitrofanovite	Pt_3Te_4	1	92
4	Moctezumite	$\text{Pb}(\text{UO}_2)(\text{TeO}_3)_2$	1	93
5	Mojaveite	$\text{Cu}_6[\text{TeO}_4(\text{OH})_2](\text{OH})_7\text{Cl}$	9	94
2	Moncheite	PtTe_2	133	70
3	Monchetundraite	Pd_2NiTe_2	1	95
4	Montanite	$\text{Bi}_2(\text{TeO}_6) \cdot 2\text{H}_2\text{O}$	22	96
2	Montbrayite	Au_2Te_3	19	97
4	Mroseite	$\text{CaTeO}_2(\text{CO}_3)$	3	98
5	Müllerite	$\text{Pb}_2\text{Fe}(\text{TeO}_6)\text{Cl}$	1	99
3	Museumite	$\text{Pb}_2\text{Pb}_2\text{S}_8\text{Te}_2$	1	100
3	Muthmannite	AuAgTe_2	7	101
6	Nabokoite	$\text{Cu}_7\text{TeO}_4(\text{SO}_4)_5 \cdot \text{KCl}$	1	102
4	Northstarite	$\text{Pb}_6(\text{TeO}_3)_5(\text{S}^{6+}\text{O}_3\text{S}^{2-})$	1	103
3	Ognitite	NiBiTe	1	104
3	Ottoite	Pb_2TeO_5	5	105
5	Oulankaite	$\text{Pd}_5\text{Cu}_4\text{SnTe}_2\text{S}_2$	3	106
3	Pampaloite	AuSbTe	1	107
5	Pararaisaite	$\text{CuMg}[\text{TeO}_4(\text{OH})_2] \cdot 6\text{H}_2\text{O}$	1	108
2	Paratellurite	TeO_2	20	109
5	Paratimroseite	$\text{Pb}_2\text{Cu}_4(\text{TeO}_6)_2 \cdot 2\text{H}_2\text{O}$	2	110
3	Pašavaite	$\text{Pd}_3\text{Pb}_2\text{Te}_2$	1	111
3	Petzite	Ag_3AuTe_2	409	112
2	Pilsenite	Bi_4Te_3	57	113
3	Pingguite	$\text{Bi}_6\text{Te}_2\text{O}_{13}$	5	114
3	Plumbotellurite	$\text{Pb}(\text{TeO}_3)$	3	115
5	Poughite	$\text{Fe}_2(\text{TeO}_3)_2(\text{SO}_4) \cdot 3\text{H}_2\text{O}$	12	116

5	Quetzalcoatlite	$\text{Cu}_4\text{Zn}_8(\text{TeO}_3)_3(\text{OH})_{18}$	6	117
4	Radhakrishnaite	PbTe_3Cl_2	2	68
5	Raisaite	$\text{CuMg}[\text{TeO}_4(\text{OH})_2] \cdot 6\text{H}_2\text{O}$	1	118
3	Rajite	CuTe_2O_5	2	119
2	Rickardite	Cu_3Te_2	52	120
5	Rodalquilarite	$\text{H}_3\text{Fe}_2(\text{Te}^{4+}\text{O}_3)_4\text{Cl}$	6	121
3	Rucklidgeite	PbBi_2Te_4	64	122
4	Saddlebackite	$\text{Pb}_2\text{Bi}_2\text{Te}_2\text{S}_3$	3	123
5	Schieffelinite	$\text{Pb}_{10}\text{Te}_6\text{O}_{20}(\text{OH})_{14}(\text{SO}_4) \cdot 5\text{H}_2\text{O}$	4	124
3	Schmitterite	$(\text{UO}_2)(\text{TeO}_3)$	6	125
2	Shuangfengite	IrTe_2	4	126
3	Skippenite	$\text{Bi}_2\text{Se}_2\text{Te}$	6	127
3	Smirnite	Bi_2TeO_5	7	128
4	Sonoraite	$\text{Fe}(\text{TeO}_3)(\text{OH}) \cdot \text{H}_2\text{O}$	11	129
3	Sopcheite	$\text{Ag}_4\text{Pd}_3\text{Te}_4$	24	130
2	Spiridonovite	Cu_2Te	1	131
3	Spiroffite	$\text{Mn}_2\text{Te}_3\text{O}_8$	4	132
2	Stützite	Ag_5Te_3	59	133
3	Sulphotsumoite	$\text{Bi}_3\text{Te}_2\text{S}$	13	134
3	Sylvanite	AgAuTe_4	337	135
5	Tamboite	$\text{Fe}_3(\text{OH})(\text{SO}_4)(\text{Te}^{4+}\text{O}_3)_3[\text{Te}^{4+}\text{O}(\text{OH})_2] \cdot 5\text{H}_2\text{O}$	1	88
4	Teineite	$\text{Cu}(\text{TeO}_3) \cdot 2\text{H}_2\text{O}$	14	136
2	Telargpalite	Pd_3Te	18	137
2	Tellurantimony	Sb_2Te_3	29	82
2	Tellurite	TeO_2	52	138
1	Tellurium	Te	190	139
2	Tellurobismuthite	Bi_2Te_3	276	140
4	Tellurohauchecornite	$\text{Ni}_9\text{BiTeS}_8$	2	141
4	Telluromandarinoite	$\text{Fe}^{3+}_2(\text{Te}^{4+}\text{O}_3)_3 \cdot 6\text{H}_2\text{O}$	1	142
3	Telluronevskite	Bi_3TeSe_2	3	143
2	Telluropalladinite	Pd_9Te_4	15	62
4	Telluroperite	$\text{Pb}_2(\text{TePb})_{\Sigma 2}\text{O}_4\text{Cl}_2$	4	144
3	Temagamite	Pd_3HgTe_3	24	145
3	Tetradymite	$\text{Bi}_2\text{Te}_2\text{S}$	510	146
4	Tewite	$\text{K}_4(\text{Te}_3\Box)_{\Sigma 4}\text{W}_{10}\text{O}_{38}$	1	147
6	Thorneite	$\text{Pb}_6(\text{Te}_2\text{O}_{10})(\text{CO}_3)\text{Cl}_2(\text{H}_2\text{O})$	4	148
5	Timroseite	$\text{Pb}_2\text{Cu}_5(\text{TeO}_6)_2(\text{OH})_2$	8	110
6	Tlalocite	$\text{Cu}_{10}\text{Zn}_6(\text{Te}^{4+}\text{O}_3)(\text{Te}^{6+}\text{O}_4)_2\text{Cl}(\text{OH})_{25} \cdot 27\text{H}_2\text{O}$	2	149
6	Tlapallite	$\text{H}_6\text{Ca}_2\text{Cu}_3\text{O}_2(\text{SO}_4)(\text{Te}^{4+}\text{O}_3)_4(\text{Te}^{6+}\text{O}_4)$	5	150
3	Törnroosite	$\text{Pd}_{11}\text{As}_2\text{Te}_2$	4	151
4	Tsnigriite	$\text{Ag}_9\text{SbTe}_3\text{S}_3$	4	152

2	Tsumoite	BiTe	100	153
5	Utahite	$\text{Cu}_5\text{Zn}_3(\text{TeO}_4)_4(\text{OH})_8 \cdot 7\text{H}_2\text{O}$	4	154
3	Vavřinite	Ni_2SbTe_2	2	155
3	Vihorlatite	$\text{Bi}_{24}\text{Se}_{17}\text{Te}_4$	1	156
3	Volynskite	AgBiTe_2	64	157
2	Vulcanite	CuTe	11	158
3	Walfordite	$(\text{Fe}^{3+}_2\text{Te}^{6+})_{\Sigma 3}\text{Te}^{4+}_9\text{O}_{24}$	1	159
2	Weissite	Cu_2Te	31	160
3	Winstanleyite	TiTe_3O_8	2	42
5	Xocolatlite	$\text{Ca}_2\text{Mn}_2\text{Te}_2\text{O}_{12} \cdot \text{H}_2\text{O}$	2	161
4	Xocomecatlite	$\text{Cu}(\text{TeO}_4)(\text{OH})_4$	10	149
4	Yafsoanite	$\text{Ca}_3\text{Zn}_3(\text{TeO}_6)_2$	4	162
5	Yecoraite	$\text{Fe}_3\text{Bi}_5\text{O}_9(\text{Te}^{4+}\text{O}_3)(\text{Te}^{6+}\text{O}_4)_2 \cdot 9\text{H}_2\text{O}$	7	163
6	Zemannite	$\text{MgZn}_2\text{Fe}_2(\text{TeO}_3)_6 \cdot 9\text{H}_2\text{O}$	7	164
3	Zincospioffite	$\text{Zn}_2\text{Te}_3\text{O}_8$	4	165

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Table S2. Classification of mineral systems of tellurium minerals according to number of species-defining elements (N) and chemical composition.

N	System	Mineral	Chemical formula
Native elements			
1	Te	Tellurium	Te
Oxides			
2	OTe	Paratellurite	TeO ₂
		Tellurite	TeO ₂
Tellurides			
<i>I. Simple tellurides</i>			
	TeAs	Kalgoorlieite	As ₂ Te ₃
	TeSb	Tellurantimony	Sb ₂ Te ₃
	TeBi	Tsumoite	BiTe
		Tellurobismuthite	Bi ₂ Te ₃
		Pilsenite	Bi ₄ Te ₃
		Hedleyite	Bi ₇ Te ₃
	TePb	Altaite	PbTe
	TeFe	Frohbergite	FeTe ₂
	TeCo	Mattagamite	CoTe ₂
	TeNi	Melonite	NiTe ₂
2	TePd	Merenskyite	PdTe ₂
		Kotulskite	PdTe
		Telargpalite	Pd ₃ Te
		Telluropalladinite	Pd ₉ Te ₄
		Keithconnite	Pd ₂₀ Te ₇
	TeIr	Shuangfengite	IrTe ₂
		Gaotaiite	Ir ₃ Te ₈
	TePt	Moncheite	PtTe ₂
		Mitrofanovite	Pt ₃ Te ₄
	TeCu	Vulcanite	CuTe
Weissite		Cu _{2-x} Te	
Rickardite		Cu _{3-x} Te ₂	
Cameronite		Cu ₅ Cu ₃ Te ₁₀	
Spiridonovite		Cu ₂ Te	
TeAg	Empressite	AgTe	
	Hessite	Ag ₂ Te	
	Stützite	Ag ₅ Te ₃	
	Lingbaoite	AgTe ₃	
TeAu	Calaverite	AuTe ₂	

		Montbrayite	Au_2Te_3
	TeHg	Coloradoite	HgTe
	TeAsPd	Törnroosite	$\text{Pd}_{11}\text{As}_2\text{Te}_2$
	TeSbNi	Vavřínite	Ni_2SbTe_2
	TeSbPd	Borovskite	Pd_3SbTe_4
	TeSbAu	Pampaloite	AuSbTe
	TeBiPb	Rucklidgeite	PbBi_2Te_4
		Kochkarite	PbBi_4Te_7
	TeBiNi	Ognitite	NiBiTe
	TeBiPd	Michenerite	PdBiTe
	TeBiIr	Mayingite	IrBiTe
	TeBiPt	Maslovite	PtBiTe
	TeBiAg	Volynskite	AgBiTe_2
	TeSnPd	Kojonenite	Pd_6SnTe_2
3	TePbPd	Pašavaite	$\text{Pd}_3\text{Pb}_2\text{Te}_2$
	TeTlAu	Honeaite	Au_3TlTe_2
	TeNiPd	Monchetundraite	Pd_2NiTe_2
	TePdAg	Lukkulaisvaaraite	$\text{Ag}_2\text{Pd}_{14}\text{Te}_9$
		Sopcheite	$\text{Ag}_4\text{Pd}_3\text{Te}_4$
	TePdHg	Temagamite	Pd_3HgTe_3
	TeCuAg	Henryite	$\text{Cu}_4\text{Ag}_3\text{Te}_4$
	TeCuAu	Kostovite	AuCuTe_4
		Bezsmertnovite	Au_4CuTe
		Muthmannite	AuAgTe_2
	TeAgAu	Petzite	AuAg_3Te_2
		Sylvanite	AuAgTe_4
		Krennerite	Au_3AgTe_8
4	TePbCuAu	Bilibinskite	$\text{PbAu}_3\text{Cu}_2\text{Te}_2$
5	TeSbPbAgHg	Mazzettiite	$\text{Ag}_3\text{HgPbSbTe}_5$

II. Sulphido-tellurides

		Ingodite	Bi_2TeS
	STeBi	Tetradymite	$\text{Bi}_2\text{Te}_2\text{S}$
		Sulphotsumoite	$\text{Bi}_3\text{Te}_2\text{S}$
3		Baksanite	$\text{Bi}_6\text{Te}_2\text{S}_3$
	STePb	Museumite	$\text{Pb}_2\text{Pb}_2\text{S}_8\text{Te}_2$
	STeCu	Goldfieldite	$\text{Cu}_{10}\text{Te}_4\text{S}_{13}$
	STeAg	Cervelleite	Ag_4TeS
4	STeSbAg	Tsnigriite	$\text{Ag}_9\text{SbTe}_3\text{S}_3$
	STeBiPb	Aleksite	$\text{PbBi}_2\text{Te}_2\text{S}_2$

		Saddlebackite	$\text{Pb}_2\text{Bi}_2\text{Te}_2\text{S}_3$
		Hitachiite	$\text{Pb}_5\text{Bi}_2\text{Te}_2\text{S}_6$
	STeBiNi	Tellurohauchecornite	$\text{Ni}_9\text{BiTeS}_8$
	STeGeAg	Alburnite	$\text{Ag}_8\text{GeTe}_2\text{S}_4$
	STeFeAg	Chenguodaite	$\text{Ag}_9\text{FeTe}_2\text{S}_4$
	STeAsAgHg	Debattistiite	$\text{Ag}_9\text{Hg}_{0.5}\text{As}_6\text{S}_{12}\text{Te}_2$
5	STeSbCuAg	Benleonardite	$\text{Ag}_{15}\text{CuSb}_2\text{S}_7\text{Te}_4$
	STeBiPbAu	Buckhornite	$\text{Pb}_2\text{AuBiTe}_2\text{S}_3$
	STeSnPdCu	Oulankaite	$\text{Pd}_5\text{Cu}_4\text{SnTe}_2\text{S}_2$
III. Selenido-tellurides			
		Kawazulite	$\text{Bi}_2\text{Te}_2\text{Se}$
	SeTeBi	Skippenite	$\text{Bi}_2\text{Se}_2\text{Te}$
		Telluronevskite	Bi_3TeSe_2
		Vihorlatite	$\text{Bi}_{24}\text{Se}_{17}\text{Te}_4$
3	SeTeNi	Kitkaite	NiTeSe
	SeTePd	Miessiite	$\text{Pd}_{11}\text{Te}_2\text{Se}_2$
	SeTeAg	Kurilite	$\text{Ag}_8\text{Te}_3\text{Se}$
	SeTeAu	Maletoyvayamite	$\text{Au}_3\text{Se}_4\text{Te}_6$
IV. Chlorido-tellurides			
3	ClTePb	Kolarite	PbTeCl_2
		Radhakrishnaite	PbTe_3Cl_2
Tellurites			
I. Anhydrous tellurites			
	OTeBi	Chekhovichite	$\text{Bi}^{3+}_2\text{Te}^{4+}_4\text{O}_{11}$
		Pingguite	$\text{Bi}_6\text{Te}_2\text{O}_{13}$
		Smirnite	Bi_2TeO_5
	OTePb	Fairbankite	$\text{Pb}(\text{TeO}_3)$
		Plumbotellurite	$\text{Pb}(\text{TeO}_3)$
	OTeTi	Winstanleyite	TiTe_3O_8
3	OTeMn	Spiroffite	$\text{Mn}_2\text{Te}_3\text{O}_8$
	OTeCu	Balyakinite	$\text{Cu}(\text{TeO}_3)$
		Rajite	CuTe_2O_5
	OTeZn	Zincospiroffite	$\text{Zn}_2\text{Te}_3\text{O}_8$
	OTeHg	Magnolite	$\text{Hg}_2(\text{TeO}_3)$
	OTeU	Cliffordite	UTe_3O_9
		Schmitterite	$(\text{UO}_2)(\text{TeO}_3)$
4	OCITePb	Telluroperite	$\text{Pb}_2(\text{TePb})_{\Sigma 2}\text{O}_4\text{Cl}_2$
	OSTeBi	Bodieite	$\text{Bi}_2(\text{TeO}_3)_2(\text{SO}_4)$

	OSTePb	Adanite	$Pb_2(TeO_3)(SO_4)$
		Northstarite	$Pb_6(Te^{4+}O_3)_5(S^{6+}O_3S^{2-})$
	OTeCCa	Mroseite	$CaTeO_2(CO_3)$
	OTeWK	Tewite	$K_4(Te_3\Box)_{\Sigma 4}W_{10}O_{38}$
	OTeMnCa	Denningite	$CaMnTe_4O_{10}$
	OTeUPb	Moctezumite	$Pb(UO_2)(TeO_3)_2$
	OCISTePbFe	Eztlite	$Pb_2Fe_3(TeO_3)_3(SO_4)O_2Cl$
6	OCISTeCuK	Nabokoite	$Cu_7TeO_4(SO_4)_5 \cdot KCl$
	OCITeSbPbCu	Choloalite	$Pb_3(Cu_2Sb)_{\Sigma 3}Te_6O_{18}Cl$

II. *Hydrous tellurites*

		Emmonsite	$Fe_2(TeO_3)_3 \cdot 2H_2O$
		Mackayite	$FeTe_2O_5(OH)$
	OHTeFe	Sonoraite	$Fe(TeO_3)(OH) \cdot H_2O$
4		Telluromandarinoite	$Fe_2(TeO_3)_3 \cdot 6H_2O$
		Millsite	$CuTeO_3 \cdot 2H_2O$
	OHTeCu	Teineite	$Cu(TeO_3) \cdot 2H_2O$
		Graemite	$Cu(TeO_3) \cdot H_2O$
	OHCITeFe	Rodalquilarite	$H_3Fe_2(TeO_3)_4Cl$
		Tamboite	$Fe_3(TeO_3)_3[TeO(OH)_2](SO_4)(OH) \cdot 5H_2O$
5	OHSTeFe	Metatamboite	$Fe_3(TeO_3)_3[TeO(OH)_2](SO_4)(OH) \cdot 3H_2O$
		Poughite	$Fe_2(TeO_3)_2(SO_4) \cdot 3H_2O$
	OHCITeAsCu	Eurekadumpite	$Cu_{16}(TeO_3)_2(AsO_4)_3Cl(OH)_{18} \cdot 7H_2O$
	OHTeAsCuCa	Juabite	$CaCu_{10}(TeO_3)_4(AsO_4)_4(OH)_2 \cdot 4H_2O$
	OHTeMnFeMg	Kinichilite	$MgMn_2Fe_2(TeO_3)_6 \cdot 9H_2O$
6	OHTeMnZnMg	Iirneyite	$MgZn_2Mn_2(TeO_3)_6 \cdot 9H_2O$
	OHTeFeNiMg	Keystoneite	$MgNi_2Fe_2(TeO_3)_6 \cdot 9H_2O$
	OHTeFeZnMg	Zemannite	$MgZn_2Fe_2(TeO_3)_6 \cdot 9H_2O$

Tellurates

I. *Anhydrous tellurates*

	OTePb	Ottoite	$Pb_2(TeO_5)$
3	OTeFe	Andymcdonaldite	$Fe_2(TeO_6)$
	OTeCu	Mcalpineite	$Cu_3(TeO_6)$
	OTeZn	Dagenaisite	$Zn_3(TeO_6)$
	OTeBiMo	Chiluite	$Bi_6Te_2Mo_2O_{21}$
4	OTePbU	Markcooperite	$Pb_2(UO_2)(TeO_6)$
	OTeMnPb	Kuranakhite	$PbMn(TeO_6)$
	OTeZnCa	Yafsoanite	$Ca_3Zn_3(TeO_6)_2$
5	OCITePbAl	Backite	$Pb_2Al(TeO_6)Cl$

	OCITePbFe	Müllerite	$\text{Pb}_2\text{Fe}(\text{TeO}_6)\text{Cl}$
	OTeAsPbZn	Dugganite	$\text{Pb}_3\text{Zn}_3(\text{TeO}_6)(\text{AsO}_4)_2$
	OTePPbZn	Kuksite	$\text{Pb}_3\text{Zn}_3(\text{TeO}_6)(\text{PO}_4)_2$
	OTePbVZn	Cheremnykhite	$\text{Pb}_3\text{Zn}_3(\text{TeO}_6)(\text{VO}_4)_2$
6	OCITeCPbCu	Hagstromite	$\text{Pb}_8\text{Cu}(\text{TeO}_6)_2(\text{CO}_3)\text{Cl}_4$
	OTeSiPbAlFe	Burckhardtite	$\text{Pb}_2(\text{Fe}^{3+}\text{Te}^{6+})_{\Sigma 2}(\text{AlSi}_3\text{O}_8)\text{O}_6$

II. Hydrous tellurates

	OHTeBi	Montanite	$\text{Bi}_2(\text{TeO}_6) \cdot 2\text{H}_2\text{O}$
	OHTeFe	Cuzticite	$\text{Fe}_2(\text{TeO}_6) \cdot 3\text{H}_2\text{O}$
4		Brumadoite	$\text{Cu}_3(\text{TeO}_4)(\text{OH})_4 \cdot 5\text{H}_2\text{O}$
		Cesbronite	$\text{Cu}_3(\text{TeO}_4)(\text{OH})_4$
	OHTeCu	Frankhawthorneite	$\text{Cu}_2(\text{TeO}_4)(\text{OH})_2$
		Jensenite	$\text{Cu}_3(\text{TeO}_6) \cdot 2\text{H}_2\text{O}$
		Xocomecatlite	$\text{Cu}_3(\text{TeO}_4)(\text{OH})_4$
	OHCITeCu	Mojaveite	$\text{Cu}_6[\text{TeO}_4(\text{OH})_2](\text{OH})_7\text{Cl}$
	OHSTePb	Schieffelinite	$\text{Pb}_{10}\text{Te}_6\text{O}_{20}(\text{OH})_{14}(\text{SO}_4) \cdot 5\text{H}_2\text{O}$
	OHTePbCr	Chromschieffelinite	$\text{Pb}_{10}\text{Te}_6\text{O}_{20}(\text{OH})_{14}(\text{CrO}_4) \cdot 5\text{H}_2\text{O}$
		Andychristyite	$\text{PbCuTeO}_5 \cdot \text{H}_2\text{O}$
		Housleyite	$\text{Pb}_6\text{CuTe}_4\text{O}_{18}(\text{OH})_2$
	OHTePbCu	Khinite	$\text{Cu}_3\text{Pb}(\text{TeO}_6)(\text{OH})_2$
		Paratimroseite	$\text{Pb}_2\text{Cu}_4(\text{TeO}_6)_2 \cdot 2\text{H}_2\text{O}$
5		Timroseite	$\text{Pb}_2\text{Cu}_5(\text{TeO}_6)_2(\text{OH})_2$
	OHTeMnCa	Xocolatlite	$\text{Ca}_2\text{Mn}_2\text{Te}_2\text{O}_{12} \cdot \text{H}_2\text{O}$
	OHTeCuZn	Utahite	$\text{Cu}_5\text{Zn}_3(\text{TeO}_4)_4(\text{OH})_8 \cdot 7\text{H}_2\text{O}$
		Quetzalcoatlite	$\text{Cu}_4\text{Zn}_8(\text{TeO}_3)_3(\text{OH})_{18}$
		Leisingite	$\text{CuMg}_2(\text{TeO}_6) \cdot 6\text{H}_2\text{O}$
	OHTeCuMg	Pararaisaite	$\text{CuMg}[\text{TeO}_4(\text{OH})_2] \cdot 6\text{H}_2\text{O}$
		Raisaite	$\text{CuMg}[\text{TeO}_4(\text{OH})_2] \cdot 6\text{H}_2\text{O}$
	OHTeCuCa	Eckhardite	$\text{CaCuTeO}_5 \cdot \text{H}_2\text{O}$
	OHCITeCPb	Thorneite	$\text{Pb}_6(\text{Te}_2\text{O}_{10})(\text{CO}_3)\text{Cl}_2 \cdot \text{H}_2\text{O}$
6	OHCITePbCu	Fuettererite	$\text{Pb}_3\text{Cu}_6(\text{TeO}_6)(\text{OH})_7\text{Cl}_5$
	OHSTePbCu	Bairdite	$\text{Pb}_2\text{Cu}_4\text{Te}_2\text{O}_{10}(\text{OH})_2(\text{SO}_4) \cdot \text{H}_2\text{O}$
	OHTeCPbCu	Agaitite	$\text{Pb}_3\text{CuTeO}_5(\text{OH})_2(\text{CO}_3)$

Tellurito-tellurates

I. Anhydrous tellurito-tellurates

3	OTeCa	Carlfriesite	$\text{CaTe}(\text{Te}^{4+})_2\text{O}_8$
	OTeFe	Walfordite	$(\text{Fe}^{3+}_2\text{Te}^{6+})_{\Sigma 3}\text{Te}^{4+}_9\text{O}_{24}$

II. Hydrous tellurito-tellurates

5	OHTeBiFe	Yecoraite	$\text{Fe}_3\text{Bi}_5\text{O}_9(\text{Te}^{4+}\text{O}_3)(\text{Te}^{6+}\text{O}_4)_2 \cdot 9\text{H}_2\text{O}$
6	OHCITeCuZn	Tlalocite	$\text{Cu}_{10}\text{Zn}_6(\text{Te}^{4+}\text{O}_3)(\text{Te}^{6+}\text{O}_4)_2\text{Cl}(\text{OH})_{25} \cdot 27\text{H}_2\text{O}$
	OHSTeCuCa	Tlapallite	$\text{H}_6\text{Ca}_2\text{Cu}_3\text{O}_2(\text{SO}_4)(\text{Te}^{4+}\text{O}_3)_4(\text{Te}^{6+}\text{O}_4)$
